

Example Items

Chemistry

Chemistry Example Items are a **representative set** of items for the ACP. Teachers may use this set of items along with the test blueprint as guides to prepare students for the ACP. On the last page, the correct answer, content SE and SE justification are listed for each item.

*The specific part of an SE that an Example Item measures is **NOT** necessarily the only part of the SE that is assessed on the ACP.* None of these Example Items will appear on the ACP.

Teachers may provide feedback regarding Example Items.

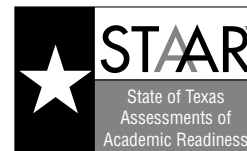
(1) Download the [Example Feedback Form](#) and email it. The form is located on the homepage of Assessment.dallasisd.org.

OR

(2) To submit directly, click “Example Feedback” **after** you login to the [Assessment website](#).

First Semester
2017–2018
Code #: 3101

STAAR CHEMISTRY REFERENCE MATERIALS



ATOMIC STRUCTURE

Speed of light = (frequency)(wavelength)

$$c = f\lambda$$

Energy = (Planck's constant)(frequency)

$$E_{\text{photon}} = hf$$

Energy = $\frac{(\text{Planck's constant})(\text{speed of light})}{(\text{wavelength})}$

$$E_{\text{photon}} = \frac{hc}{\lambda}$$

BEHAVIOR OF GASES

Total pressure of a gas = $\left(\begin{array}{l} \text{sum of the partial pressures} \\ \text{of the component gases} \end{array} \right)$

$$P_T = P_1 + P_2 + P_3 + \dots$$

(Pressure)(volume) = (moles)(ideal gas constant)(temperature)

$$PV = nRT$$

$\frac{(\text{Initial pressure})(\text{initial volume})}{(\text{Initial moles})(\text{initial temperature})} = \frac{(\text{final pressure})(\text{final volume})}{(\text{final moles})(\text{final temperature})}$

$$\frac{P_1V_1}{n_1T_1} = \frac{P_2V_2}{n_2T_2}$$

(Initial pressure)(initial volume) = (final pressure)(final volume)

$$P_1V_1 = P_2V_2$$

$\frac{(\text{Initial volume})}{(\text{Initial temperature})} = \frac{(\text{final volume})}{(\text{final temperature})}$

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

$\frac{(\text{Initial volume})}{(\text{Initial moles})} = \frac{(\text{final volume})}{(\text{final moles})}$

$$\frac{V_1}{n_1} = \frac{V_2}{n_2}$$

SOLUTIONS

Molarity = $\frac{\text{moles of solute}}{\text{liter of solution}}$

$$M = \frac{\text{mol}}{\text{L}}$$

Ionization constant of water = $\left(\begin{array}{l} \text{hydrogen ion} \\ \text{concentration} \end{array} \right) \left(\begin{array}{l} \text{hydroxide ion} \\ \text{concentration} \end{array} \right)$

$$K_w = [\text{H}^+][\text{OH}^-]$$

$\left(\begin{array}{l} \text{Volume of} \\ \text{solution 1} \end{array} \right) \left(\begin{array}{l} \text{molarity of} \\ \text{solution 1} \end{array} \right) = \left(\begin{array}{l} \text{volume of} \\ \text{solution 2} \end{array} \right) \left(\begin{array}{l} \text{molarity of} \\ \text{solution 2} \end{array} \right)$

$$V_1M_1 = V_2M_2$$

pH = -logarithm (hydrogen ion concentration)

$$\text{pH} = -\log[\text{H}^+]$$

THERMOCHEMISTRY

Heat gained or lost = (mass) $\left(\begin{array}{l} \text{specific} \\ \text{heat} \end{array} \right) \left(\begin{array}{l} \text{change in} \\ \text{temperature} \end{array} \right)$

$$Q = mc_p\Delta T$$

Enthalpy of reaction = $\left(\begin{array}{l} \text{enthalpy} \\ \text{of products} \end{array} \right) - \left(\begin{array}{l} \text{enthalpy} \\ \text{of reactants} \end{array} \right)$

$$\Delta H = \Delta H_f^\circ(\text{products}) - \Delta H_f^\circ(\text{reactants})$$

STAAR CHEMISTRY REFERENCE MATERIALS

OTHER FORMULAS

$$\text{Density} = \frac{\text{mass}}{\text{volume}}$$

$$D = \frac{m}{V}$$

$$\text{Percent error} = \left(\frac{\text{accepted value} - \text{experimental value}}{\text{accepted value}} \right) (100)$$

$$\text{Percent yield} = \left(\frac{\text{actual yield}}{\text{theoretical yield}} \right) (100)$$

CONSTANTS AND CONVERSIONS

$$\text{Avogadro's number} = 6.02 \times 10^{23} \text{ particles per mole}$$

$$h = \text{Planck's constant} = 6.63 \times 10^{-34} \text{ J} \cdot \text{s}$$

$$c = \text{speed of light} = 3.00 \times 10^8 \frac{\text{m}}{\text{s}}$$

$$K_w = \text{ionization constant of water} = 1.00 \times 10^{-14} \left(\frac{\text{mol}}{\text{L}} \right)^2$$

$$\text{alpha particle } (\alpha) = {}_2^4\text{He} \quad \text{beta particle } (\beta) = {}_{-1}^0\text{e} \quad \text{neutron} = {}_0^1\text{n}$$

$$\text{standard temperature and pressure (STP)} = 0^\circ\text{C and 1 atm}$$

$$0^\circ\text{C} = 273 \text{ K}$$

$$\text{volume of ideal gas at STP} = 22.4 \frac{\text{L}}{\text{mol}}$$

$$1 \text{ cm}^3 = 1 \text{ mL} = 1 \text{ cc}$$

$$1 \text{ atm} = 760 \text{ mm Hg} = 101.3 \text{ kPa}$$

$$R = \text{ideal gas constant} = 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} = 8.31 \frac{\text{L} \cdot \text{kPa}}{\text{mol} \cdot \text{K}} = 62.4 \frac{\text{L} \cdot \text{mm Hg}}{\text{mol} \cdot \text{K}}$$

$$1 \text{ calorie (cal)} = 4.18 \text{ joules (J)}$$

$$1000 \text{ calories (cal)} = 1 \text{ Calorie (Cal)} = 1 \text{ kilocalorie (kcal)}$$

RULES FOR SIGNIFICANT FIGURES

1. Non-zero digits and zeros between non-zero digits are always significant.
2. Leading zeros are not significant.
3. Zeros to the right of all non-zero digits are only significant if a decimal point is shown.
4. For values written in scientific notation, the digits in the coefficient are significant.
5. In a common logarithm, there are as many digits after the decimal point as there are significant figures in the original number.

STAAR CHEMISTRY REFERENCE MATERIALS

POLYATOMIC IONS		SOLUBILITY OF COMMON IONIC COMPOUNDS IN WATER		ACTIVITY SERIES
Acetate	$C_2H_3O_2^-$, CH_3COO^-	<u>Soluble</u> <u>compounds contain</u> $C_2H_3O_2^-$, CH_3COO^-	<u>Common exceptions</u> None	<u>Metal</u> Lithium
Ammonium	NH_4^+	NH_4^+	None	Potassium
Carbonate	CO_3^{2-}	CO_3^{2-}	None	Barium
Chlorate	ClO_3^-	ClO_3^-	None	Calcium
Chlorite	ClO_2^-	ClO_2^-	None	Sodium
Chromate	CrO_4^{2-}	CrO_4^{2-}	None	Magnesium
Cyanide	CN^-	CN^-	None	Aluminum
Dichromate	$Cr_2O_7^{2-}$	$Cr_2O_7^{2-}$	None	Manganese
Hydrogen carbonate	HCO_3^-	HCO_3^-	Compounds of Ag^+ , Pb^{2+} , and Hg_2^{2+}	Zinc
Hydroxide	OH^-	OH^-	Compounds of Ag^+ , Pb^{2+} , and Hg_2^{2+}	Chromium
Hypochlorite	ClO^-	ClO^-	Compounds of Ag^+ , Pb^{2+} , and Hg_2^{2+}	Iron
Nitrate	NO_3^-	NO_3^-	Compounds of Sr^{2+} , Ba^{2+} , Pb^{2+} , and Hg_2^{2+}	Cobalt
Nitrite	NO_2^-	NO_2^-		Nickel
Perchlorate	ClO_4^-	ClO_4^-		Tin
Permanganate	MnO_4^-	$Cr_2O_7^{2-}$		Lead
Phosphate	PO_4^{3-}	OH^-		(Hydrogen)
Sulfate	SO_4^{2-}	S^{2-}		Copper
Sulfite	SO_3^{2-}			Mercury
				Silver
				Platinum
				Gold



STAAR CHEMISTRY REFERENCE MATERIALS

PERIODIC TABLE OF THE ELEMENTS

1 1A	2 2A	3 3B	4 4B	5 5B	6 6B	7 7B	8 8B	9	10	11 1B	12 2B	13 3A	14 4A	15 5A	16 6A	17 7A	18 8A														
1 H 1.008 Hydrogen	2 He 4.003 Helium	3 Li 6.941 Lithium	4 Be 9.012 Beryllium	5 B 10.812 Boron	6 C 12.011 Carbon	7 N 14.007 Nitrogen	8 O 15.999 Oxygen	9 F 18.998 Fluorine	10 Ne 20.180 Neon	11 Na 22.990 Sodium	12 Mg 24.305 Magnesium	13 Al 26.982 Aluminum	14 Si 28.086 Silicon	15 P 30.974 Phosphorus	16 S 32.066 Sulfur	17 Cl 35.453 Chlorine	18 Ar 39.948 Argon														
19 K 39.098 Potassium	20 Ca 40.078 Calcium	21 Sc 44.956 Scandium	22 Ti 47.867 Titanium	23 V 50.942 Vanadium	24 Cr 51.996 Chromium	25 Mn 54.938 Manganese	26 Fe 55.845 Iron	27 Co 58.933 Cobalt	28 Ni 58.693 Nickel	29 Cu 63.546 Copper	30 Zn 65.38 Zinc	31 Ga 69.723 Gallium	32 Ge 72.64 Germanium	33 As 74.922 Arsenic	34 Se 78.96 Selenium	35 Br 79.904 Bromine	36 Kr 83.798 Krypton														
37 Rb 85.468 Rubidium	38 Sr 87.62 Strontium	39 Y 88.906 Yttrium	40 Zr 91.224 Zirconium	41 Nb 92.906 Niobium	42 Mo 95.96 Molybdenum	43 Tc (98) Technetium	44 Ru 101.07 Ruthenium	45 Rh 102.906 Rhodium	46 Pd 106.42 Palladium	47 Ag 107.868 Silver	48 Cd 112.412 Cadmium	49 In 114.818 Indium	50 Sn 118.711 Tin	51 Sb 121.760 Antimony	52 Te 127.60 Tellurium	53 I 126.904 Iodine	54 Xe 131.294 Xenon														
55 Cs 132.905 Cesium	56 Ba 137.328 Barium	57 La 138.905 Lanthanum	58 Ce 140.116 Cerium	59 Pr 140.908 Praseodymium	60 Nd 144.242 Neodymium	61 Pm (145) Promethium	62 Sm 150.36 Samarium	63 Eu 151.964 Europium	64 Gd 157.25 Gadolinium	65 Tb 158.925 Terbium	66 Dy 162.500 Dysprosium	67 Ho 164.930 Holmium	68 Er 167.259 Erbium	69 Tm 168.934 Thulium	70 Yb 173.055 Ytterbium	71 Lu 174.967 Lutetium	72 Hf 178.49 Hafnium	73 Ta 180.948 Tantalum	74 W 183.84 Tungsten	75 Re 186.207 Rhenium	76 Os 190.23 Osmium	77 Ir 192.217 Iridium	78 Pt 195.085 Platinum	79 Au 196.967 Gold	80 Hg 200.59 Mercury	81 Tl 204.383 Thallium	82 Pb 207.2 Lead	83 Bi 208.980 Bismuth	84 Po (209) Polonium	85 At (210) Astatine	86 Rn (222) Radon
87 Fr (223) Francium	88 Ra (226) Radium	89 Ac (227) Actinium	90 Th 232.038 Thorium	91 Pa 231.036 Protactinium	92 U 238.029 Uranium	93 Np (237) Neptunium	94 Pu (244) Plutonium	95 Am (243) Americium	96 Cm (247) Curium	97 Bk (247) Berkelium	98 Cf (251) Californium	99 Es (252) Einsteinium	100 Fm (257) Fermium	101 Md (258) Mendelevium	102 No (259) Nobelium	103 Lr (262) Lawrencium	104 Rf (267) Rutherfordium	105 Db (268) Dubnium	106 Sg (271) Seaborgium	107 Bh (272) Bohrium	108 Hs (270) Hassium	109 Mt (276) Meitnerium	110 Ds (281) Darmstadtium	111 Rg (280) Roentgenium	112 Cn (285) Copernicium	113 Nh (286) Nihonium	114 Fl (289) Flerovium	115 Mc (288) Moscovium	116 Lv (293) Livermorium	117 Ts (294) Tennessine	118 Og (294) Oganesson

Mass numbers in parentheses are those of the most stable or most common isotope.

Lanthanide Series

Actinide Series

EXAMPLE ITEMS Chemistry, Sem 1

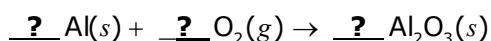
1 Which is a chemical property of copper?

- A It is a mineral.
- B It reacts with nitric acid.
- C It is a liquid at 1083 °C.
- D It is malleable and ductile.

2 What is the number of atoms in three grams of gold?

- A 9.17×10^{21}
- B 1.806×10^{24}
- C 6.02×10^{23}
- D 2.29×10^{22}

3 The unbalanced equation for the synthesis of aluminum oxide is shown.



Which list shows the order of coefficients needed to balance this equation?

- A 2, 3, 1
- B 2, 1, 2
- C 4, 1, 2
- D 4, 3, 2

4 The table shows the percent composition of a compound containing hydrogen and oxygen.

Element	Percent Composition
Hydrogen	11.19%
Oxygen	88.81%

What is the empirical formula for this compound?

- A H_2O
- B H_2O_2
- C H_{11}O_6
- D $\text{H}_{11}\text{O}_{89}$

EXAMPLE ITEMS Chemistry, Sem 1

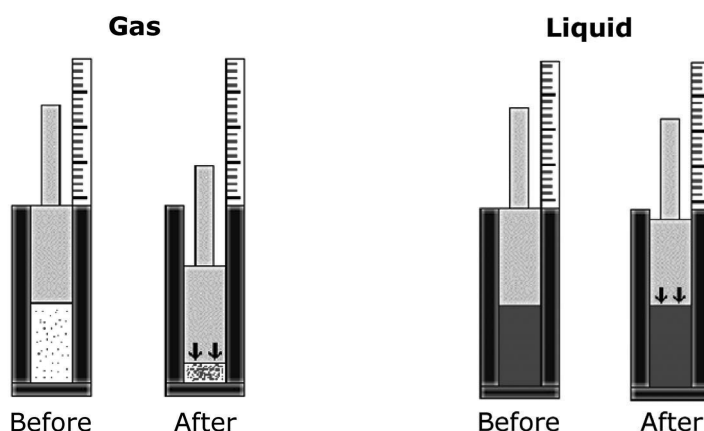
- 5** How many grams are there in one mole of Aluminum?
- A** 3.01×10^{23}
 - B** 6.02×10^{25}
 - C** 13.0
 - D** 26.982
- 6** Gold is classified as a pure substance. Which of its properties supports this classification?
- A** Gold occurs in nature as a solid.
 - B** Gold is resistant to rust and tarnish.
 - C** Gold has a high melting point.
 - D** Gold is made of one kind of atom.
- 7** What is the formula of calcium permanganate?
- A** CaMnO_4
 - B** $\text{Ca}(\text{MnO}_4)_2$
 - C** Ca_3MnO_4
 - D** $\text{Ca}_3(\text{MnO}_4)_3$
- 8** Which chemical family or group on the Periodic Table contains the least reactive elements?
- A** Group 1
 - B** Group 3
 - C** Group 17
 - D** Group 18
- 9** The Electron Sea Model explains why metals —
- A** accept electrons easily
 - B** are excellent conductors
 - C** are mostly brittle
 - D** form ionic compounds

EXAMPLE ITEMS Chemistry, Sem 1

10 What is the name of the molecular compound with the formula P_2O_5 ?

- A Pentaphosphorus dioxide
- B Decaphosphorus pentoxide
- C Diphosphorus pentoxide
- D Diphosphorus monoxide

11 Equal volumes of a liquid and a gas are placed in separate containers, and equal pressure is applied as shown in the illustration.



Which statement describes the result of this experiment?

- A The liquid has less mass than the gas.
- B The liquid has more mass than the gas.
- C The gas is less compressible than the liquid.
- D The gas is more compressible than the liquid.

12 What is the Lewis electron dot structure for the covalent compound sulfur dioxide?

- A $\ddot{O}=\ddot{S}-\ddot{O}:$
- B $\ddot{O}=\ddot{S}=\ddot{O}:$
- C $\ddot{O}-\ddot{S}-\ddot{O}:$
- D $\ddot{S}-\ddot{O}=\ddot{O}$

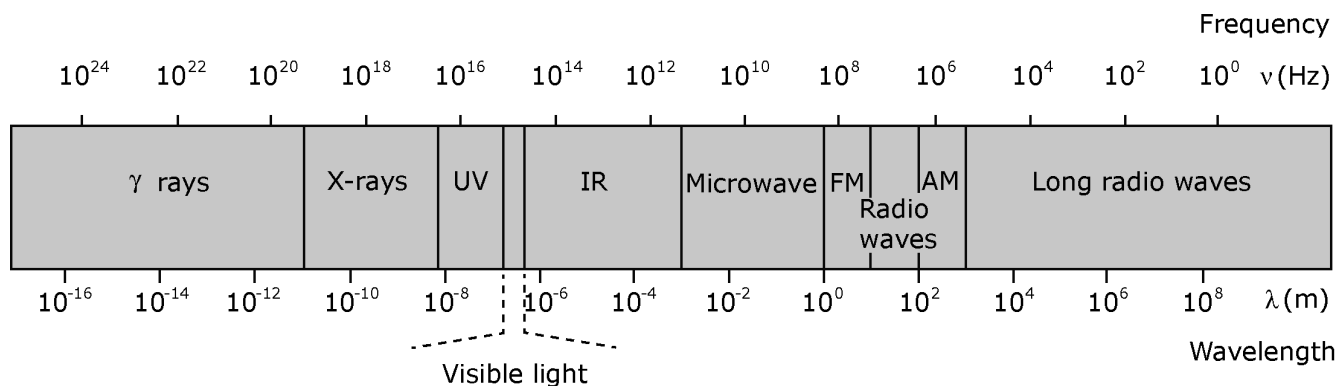
EXAMPLE ITEMS Chemistry, Sem 1

13 Which represents the ground state electron configuration for manganese?

- A $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^5$
- B $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^5$
- C $1s^2 2s^2 2p^6 3s^2 3p^6 4s^7$
- D $1s^2 2s^2 2p^6 3s^2 3p^6 4d^7$



Use the electromagnetic spectrum to answer the next question.



14 Which type of electromagnetic waves have the greatest energy?

- A Gamma rays
- B Microwaves
- C Radio waves
- D Visible light

15 If a beam of blue light has a frequency of 6.67×10^{14} Hz, its wavelength is —

- A 4.50×10^{-7} m
- B 4.50×10^{-6} m
- C 3.00×10^8 m
- D 2.22×10^6 m

EXAMPLE ITEMS Chemistry, Sem 1

16

The table shows the two isotopes of an element.

Isotope	Mass (amu)	Percent Abundance	Average Mass
x	10.013	19.8%	
y	11.009	80.2%	
Average Atomic Mass			?

What is the average atomic mass of the element?

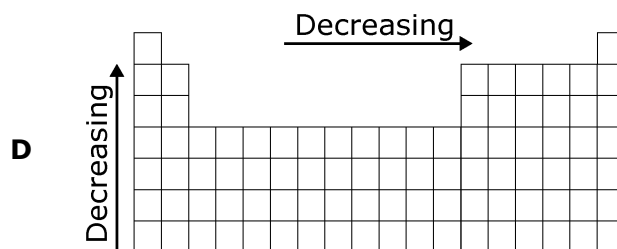
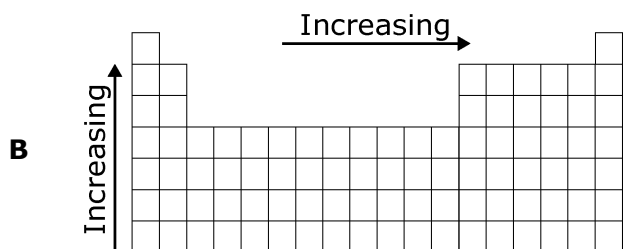
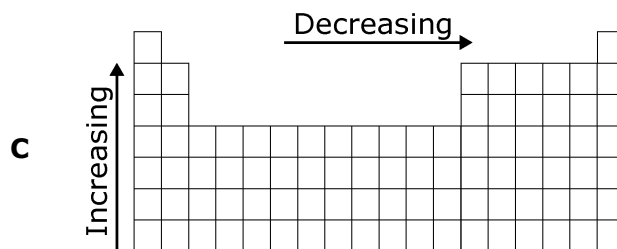
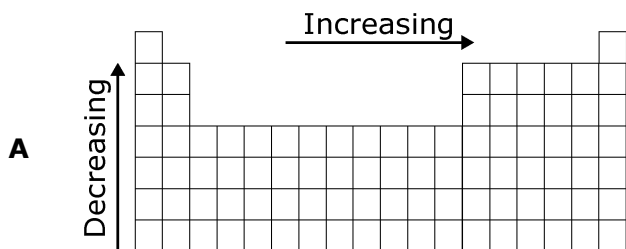
(Express the answer to three significant figures.)

Record the answer and fill in the bubbles on the grid provided. Be sure to use the correct place value.

+	•	•	•	•	•	•	•
-	0	0	0	0	0	0	0
	1	1	1	1	1	1	1
	2	2	2	2	2	2	2
	3	3	3	3	3	3	3
	4	4	4	4	4	4	4
	5	5	5	5	5	5	5
	6	6	6	6	6	6	6
	7	7	7	7	7	7	7
	8	8	8	8	8	8	8
	9	9	9	9	9	9	9

EXAMPLE ITEMS Chemistry, Sem 1

- 17 Electronegativity measures the ability of atoms to attract a bonding pair of electrons. Which diagram correctly identifies both the period and group trends for electronegativity in the Periodic Table?



EXAMPLE ITEMS Chemistry Key, Sem 1

Item#	Key	SE	Process Skills	SE Justification
1	B	CHEM.4A	--	Differentiate between physical and chemical changes and properties.
2	A	CHEM.8B	2G	Use the mole concept to calculate the number of atoms, in a sample of material.
3	D	CHEM.8D	2G	Use the law of conservation of mass to chemical equations.
4	A	CHEM.8C	2H	Calculate molecular formulas.
5	D	CHEM.8A	2G	Use the concept of a mole.
6	D	CHEM.4D	--	Classify matter as pure substances or mixtures through investigation of their properties.
7	B	CHEM.7B	--	Write the chemical formulas of common polyatomic ionic compounds containing main group or transition metals.
8	D	CHEM.5B	2H	Use the Periodic Table to identify and explain the properties of chemical families, including noble gases.
9	B	CHEM.7D	--	Apply the theory of metallic bonding to explain the electrical conductivity of metals.
10	C	CHEM.7A	--	Name covalent compounds using International Union of Pure and Applied Chemistry (IUPAC) nomenclature rules.
11	D	CHEM.4C	2H, 2I	Compare liquids, and gases in terms of compressibility.
12	B	CHEM.7C	2I	Construct electron dot formulas to illustrate covalent bonds.
13	A	CHEM.6E	--	Express the arrangement of electrons in atoms through electron configurations.
14	A	CHEM.6B	2H, 2I	Understand the electromagnetic spectrum and the mathematical relationships between energy, frequency, and wavelength of light.
15	A	CHEM.6C	2G	Calculate the wavelength of light using its frequency and the speed of light.
16	10.8	CHEM.6D	2H, 2G	Use isotopic composition to calculate average atomic mass of an element.
17	B	CHEM.5C	2H	Use the Periodic Table to identify and explain periodic trends, including electronegativity.